

B. Pharm. First Semester (New)
35302 : BP 102 T - Pharmaceutical Analysis – I

P. Pages : 2

Time : Three Hours



AU - 0701

Max. Marks : 75

- Notes : 1. Illustrate your answer necessary with the help of neat sketches.
 2. Solve All questions.

1. A) Multiple choice questions.
- 1) A solution of known concentration is the definition of 1
 a) Standard solution b) Buffer solution
 c) Neutral solution d) Saturated solution
 - 2) The indicator of Fajan's method is 1
 a) Potassium chromate b) Fluorescein
 c) Tatraine d) Both b and c
 - 3) In precipitation titration titrant is 1
 a) Sodium chloride b) Silver chloride
 c) Silver nitrate d) Sodium nitrate
 - 4) The electrochemical method which is used to measure electromotive force is called as 1
 a) Amperometry b) Conductometry
 c) Polarography d) Potentiometry
 - 5) The solution in conical flask is called as 1
 a) Titrate b) Titrant
 c) Indicator d) None of above
 - 6) The unit of conductivity is 1
 a) Ohm b) Mho
 c) Ohm meter d) Mho cm⁻¹
 - 7) The number of moles of solute per litre of solvent is known as 1
 a) Molarity b) Normality
 c) Molality d) Formality
 - 8) Example for strong acid Vs strong base titration is 1
 a) HCl Vs NaOH b) HCl Vs NH₄OH
 c) CH₃COOH Vs NaOH d) CH₃COOH Vs NH₄OH
 - 9) ----- is primary standard substance. 1
 a) Hydrochloric acid b) Potassium hydrogen phthalate
 c) Sodium hydroxide d) All of above
 - 10) The number of significant figure in 1.268g is 1
 a) 3 b) 2
 c) 4 d) 5

B) Objective type questions.

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|---|---|
| 1) Write a note on warner coordination number. | 2 |
| 2) Define accuracy and precision. | 2 |
| 3) Write note on primary and secondary standard. | 2 |
| 4) What is post precipitation. | 2 |
| 5) Give the example of electrode used in potentiometry. | 2 |

2. Long answers solve **any two**.

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| 1) Write in detail principle and steps involved in gravimetric analysis. | 10 |
| 2) Define neutralization curve and explain in detail theory involved in acid base titration. | 10 |
| 3) What is error? Explain the types and methods of minimizing error. | 10 |

3. Short Answers solve **any seven**.

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| 1) Describe in detail volhard's method & modified volhard's method. | 5 |
| 2) Write the principle and procedure for estimation of sodium chloride. | 5 |
| 3) Explain the limit test for Sulphate. | 5 |
| 4) Write a note on masking and demasking agent. | 5 |
| 5) Explain the preparation and standardization of 1 M sodium hydroxide. | 5 |
| 6) Explain construction and working of dropping mercury electrode. | 5 |
| 7) Write a note on coprecipitation and post precipitation. | 5 |
| 8) Explain the types of solvents used in nonaqueous titration. | 5 |
| 9) What are the methods of determine end point of potentiometric titration. | 5 |
