B.Pharm First Semester (New)

35302 : Pharmaceutical Analysis : BP 102 T

P. Pages : 2 | **AV - 2364**

Time : Three Hours | Max. Marks : 75

* 0 3 7 3 *

Notes : 1. Illustrate your answer necessary with the help of neat sketches.
2. Solve all questions.

1. a) Multiple choice questions

1) Which of the following is not a example for non aqueous titration. **1**
   a) Methanol b) Pyridine
   c) Dimethyl Formamide d) Water

2) Which of the following is an example for weak acid Vs strong base. **1**
   a) HCl Vs NaOH b) CH_3COOH Vs NaOH
   c) HCl Vs NH_4OH d) CH_3COOH Vs NH_4OH

3) The pH at equivalence point for strong acid & strong base is **1**
   a) 7 b) Less than 7
   c) More than 7 d) None of above

4) The number of moles of solute present in 1000g of solute is known as **1**
   a) Normality b) Molarity
   c) Molality d) Formality

5) The unit of Conductivity is **1**
   a) Ohm b) Mho
   c) Ohm meter d) Mho cm^{-1}

6) The example for Polarisable electrode is **1**
   a) Dropping mercury electrode b) Rotating platinum electrode
   c) Both a and b d) Electrochemical electrode

7) Assay of ferrous sulphate is based on -------- type of titration. **1**
   a) Aqueous b) Nonaqueous
   c) Precipitation d) Redox

8) Limit test is **1**
   a) Quantitative analysis b) Semiquantitative analysis
   c) Qualitative analysis d) Both a and b

9) The titrant used in nonaqueous titration is **1**
   a) Hydrochloric acid b) Perchloric acid
   c) Potassium Permanganate d) Sodium hydroxide

10) The sample size "Macro" for analysis is **1**
   a) 0.01 g to 0.1 g b) 0.1 g or more
   c) 10^{-4} g d) 0.001 to 0.01 g

b) Objective type questions

1) Define the term normality and molarity. **2**

2) Write a note on significant figure. **2**

3) What is levelling solvent. **2**

4) Enlist oxidizing agent and reducing agent. **2**

5) What is secondary standard and give its examples. **2**

2. Long Answer (Solve **any two**)

1) Write in detail about sources of impurities & note on limit test for chloride. **10**

2) Explain in detail the types of conductometric titrations. **10**

3) What is redox titration? Explain the types of redox titration in detail. **10**

3. Short Answer (solve **any seven**)

1) Discuss in detail Mohr's method. **5**

2) Write a note on diazotization titration. **5**

3) Define error. What are the methods of minimizing errors? **5**

4) Explain the neutralization curve for strong acid and strong base. **5**

5) Explain the concept of oxidation and reduction. **5**

6) Write a note on metal ion indicator. **5**

7) Describe the construction and working of standard hydrogen electrode. **5**

8) Write the principle and procedure for assay of ephedrine Hcl. **5**

9) Add a note on accuracy and precision. **5**

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